



Topic name	Term	Skills Developed	Link to Subject Content	Prior learning	Next curriculum link
4.1 Atomic Structure and the Periodic table	Autumn	<p>AT 4 Safe use of a range of equipment to separate chemical mixtures. WS 2.2, 2.3</p> <p>WS 1.1, 1.6 This historical context provides an opportunity for students to show an understanding of why and describe how scientific methods and theories develop over time. WS 1.2</p> <p>WS 1.2</p> <p>WS 4.3, 4 Use SI units and the prefix nano. MS 1b Recognise expressions in standard form. MS 1d</p> <p>WS 1.2 Students should be able to represent the electronic structures of the first twenty elements of the periodic table in both forms. MS 5b Visualise and represent 2D and 3D forms including two-dimensional representations of 3D objects.</p> <p>WS 1.1, 1.6 Explain how testing a prediction can support or refute a new scientific idea.</p> <p>WS 1.2</p> <p>WS 1.2</p> <p>WS 1.2 AT 6 Offers an opportunity within displacement reactions of halogens.</p>	<ul style="list-style-type: none"><li>4.1.1.2 Mixtures</li><li>4.1.1.3 The development of the model of the atom</li><li>4.1.1.4 Relative electrical charges of subatomic particles</li><li>4.1.1.5 Size and mass of atoms</li><li>4.1.1.7 Electronic structure</li><li>4.1.2.2 Development of the periodic table</li><li>4.1.2.4 Group 0</li><li>4.1.2.5 Group 1</li><li>4.1.2.6 Group 7</li></ul>	<p><i>Links from prior topics:</i></p> <p>KS3 Y7 Particles</p> <p>Y8 Atoms, Elements and Compounds</p> <p>Y8 Separating Mixtures</p>	<p><i>Links to future topics:</i></p> <p><b>GCSE:</b> Y10 - 4.7 Organic chemistry – crude oil/reactions</p> <p>Y9 - 4.2 Bonding, structure, and the properties of matter – forming ions</p> <p><b>A-level:</b> Y12 - 3.1.1 Atomic structure – development of the atom, sub atomic particles</p> <p>Y12 - 3.2.1 Periodicity</p> <p>Y12 - 3.2.2 Group 2, the alkaline earth metals</p> <p>Y12 - 3.2.2 Group 2, the alkaline earth metals</p> <p>Y13 - 3.2.5 Transition metals</p>



<p><b>4.2 Bonding, Structure and the properties of matter</b></p>	<p>Autumn</p>	<p>MS 5b Visualise and represent 2D and 3D forms including two dimensional representations of 3D objects. WS 1.2</p> <p>MS 5b Visualise and represent 2D and 3D forms including two dimensional representations of 3D objects. WS 1.2 MS 4a MS 1a, 1c</p> <p>WS 1.2 Recognise substances as small molecules, polymers or giant structures from diagrams showing their bonding. MS 5b</p> <p>WS 1.2 Recognise substances as metallic giant structures from diagrams showing their bonding. MS 5b Visualise and represent 2D and 3D forms including two dimensional representations of 3D objects.</p> <p>MS 5b Visualise and represent 2D and 3D forms including two dimensional representations of 3D objects. WS 1.2</p> <p>WS 1.2</p> <p>MS 5b Visualise and represent 2D and 3D forms including two dimensional representations of 3D objects. WS 1.2</p> <p>WS 1.2</p> <p>MS 5b Visualise and represent 2D and 3D forms including two dimensional representations of 3D objects.</p>	<ul style="list-style-type: none"> <li>● 4.2.1.2 Ionic bonding</li> <li>● 4.2.1.3 Ionic compounds</li> <li>● 4.2.1.4 Covalent bonding</li> <li>● 4.2.1.5 Metallic bonding</li> <li>● 4.2.2.1 The three states of matter</li> <li>● 4.2.2.4 Properties of small molecules</li> <li>● 4.2.2.6 Giant covalent structures</li> <li>● 4.2.2.7 Properties of metals and alloys</li> <li>● 4.2.3.1 Diamond</li> <li>● 4.2.3.2 Graphite</li> <li>● 4.2.3.3 Graphene and fullerenes</li> </ul>	<p><i>Links from prior topics:</i></p> <p>KS3 Y7 Particles</p> <p>Y8 Atoms, Element &amp; Compounds</p> <p>GCSE 4.1 Atomic structure</p>	<p><i>Links to future topics:</i></p> <p>GCSE: Y10 – 4.4.3 Electrolysis - conductivity</p> <p>A-level: Y12 - 3.1.3 Bonding</p>
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		<p>WS 1.2</p> <p>WS 1.2, 1.4 MS 5b Visualise and represent 2D and 3D forms including two dimensional representations of 3D objects.</p> <p>WS 1.3, 1.4, 1.5</p>			
<p><b>4.8</b> <b>Chemical analysis</b></p>	<p><i>Spring</i></p>	<p>MS 1a Recognise and use expressions in decimal form. MS 1c Use ratios, fractions and percentages.</p> <p>WS 1.4, 2.2, 4.1</p> <p>WS 2.2, 3.1, 2, 3</p> <p>MS 1d Make estimates of the results of simple calculations. MS 2a</p>	<ul style="list-style-type: none"> <li>• Distinguishing between pure and impure substances</li> <li>• separation techniques for mixtures of substances: filtration, crystallisation, Chromatography</li> <li>• identification of common gases</li> <li>•</li> </ul>	<p><b>KS3</b> Y8 Chemical reactions Testing for gases</p> <p>Y8 Introduction to bonding</p> <p>Y8 The Periodic table</p> <p>4.1 Atomic Structure &amp; The Periodic Table</p> <p>Group 7 Separation of mixtures</p> <p>4.2 Structure &amp; Bonding</p>	<p><i>Links to future topics:</i></p> <p><b>GCSE</b> 4.4 Chemical Changes</p> <p><b>A Level</b> <b>Group 2 and group 7 elements</b></p> <ul style="list-style-type: none"> <li>• Halide ions</li> <li>• Test for ions</li> </ul> <p>Links to GCSE Topic 7 and A' level unit 2 –</p> <p><b>Organic chemistry</b> <b>Year 11:</b></p> <ul style="list-style-type: none"> <li>• Polymers</li> </ul>
<p><b>4.4</b> <b>Chemical changes (Reactivity of metals, reactions of acids and Electrolysis)</b></p>	<p><i>Spring</i></p>	<p>AT 2 – safe use of appropriate heating devices and techniques including use of a Bunsen burner and a water bath or electric heater.</p> <p>AT 3 – use of appropriate apparatus and techniques for conducting chemical reactions, including appropriate reagents.</p>	<ul style="list-style-type: none"> <li>• 4.4.1.2 The reactivity series</li> <li>• Describe the use of universal indicator or a wide range indicator to measure the approximate pH of a solution</li> <li>• Use the pH scale to identify acidic or alkaline solutions.</li> <li>• Describe how to carry out titrations using strong acids and strong alkalis only (sulfuric,</li> </ul>	<p><b>KS3</b> Y8 Introduction to bonding</p> <p>Y8 Chemical reactions</p> <p>4.2 Structure &amp; Bonding</p>	<p><b>GCSE:</b> Y10 – 4.4.1 Reactivity of metals - extraction of metals</p> <p>Y10 - 4.4.2 Reactions of acids</p> <p><b>A-level:</b> Y13 - 3.3.9.2 Acylation (RP 10)</p>



	<p>AT 4 – safe use of a range of equipment to purify and/or separate chemical mixtures including evaporation, filtration, crystallisation.</p> <p>AT 6 – safe use and careful handling of liquids and solids, including careful mixing of reagents under controlled conditions. Key opportunities for skills development. In doing this practical there are key opportunities for students to develop the following skills.</p> <p>AT 7 – use of appropriate apparatus and techniques to draw, set up and use electrochemical cells for separation and production of elements and compounds.</p> <p>AT 8 – use of appropriate qualitative reagents and techniques to analyse and identify unknown samples or products including gas tests for hydrogen, oxygen and chlorine. Key opportunities and skills development</p> <p>WS 2.3 – apply a knowledge of a range of techniques, instruments, apparatus, and materials to select those appropriate to the experiment.</p> <p>AT 1, 3, 8 Opportunities within titrations including to determine concentrations of strong acids and alkalis.</p> <p>AT 3 This is an opportunity to investigate pH changes when a strong acid neutralises a strong alkali.</p> <p>AT 8 An opportunity to measure the pH of different acids at different concentrations.</p>	<p>hydrochloric and nitric acids only) to find the reacting volumes accurately</p> <ul style="list-style-type: none"><li>• (HT Only) Calculate the chemical quantities in titrations involving concentrations in mol/dm<sup>3</sup> and in g/dm<sup>3</sup>.</li><li>• 4.4.3.4 Electrolysis of aqueous solutions</li></ul> <p>8.2.1 Required practical activity 1 Preparation of a pure, dry sample of a soluble salt from an insoluble oxide or carbonate, using a Bunsen burner to heat dilute acid and a water bath or electric heater to evaporate the solution. Apparatus and techniques In doing this practical students should cover these parts of the apparatus and techniques requirements.</p> <p>AT skills covered by this practical activity: 1 and 8.</p> <p>8.2.3 Required practical activity 3 Investigate what happens when aqueous solutions are electrolysed using inert electrodes. This should be an investigation involving developing a hypothesis. Apparatus and techniques In doing this practical students should cover these parts of the apparatus and techniques requirements</p>	<ul style="list-style-type: none"><li>• Ionic bonding</li></ul>	<p>Y13 - 3.1.14 Electrode Potentials</p> <p>Y13 – 3.1.12 Acids &amp; Bases</p>
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<p>4.3 Quantitative Chemistry</p>	<p>Summer</p>	<p>WS 1.2 Use a variety of models such as representational, spatial, descriptive, computational and mathematical to solve problems, make predictions and to develop scientific explanations and understanding of familiar and unfamiliar facts.</p> <p>WS 3.4 Representing distributions of results and make estimations of uncertainty.</p> <p>WS 4.1 Use scientific vocabulary, terminology and definitions.</p> <p>WS 4.2 Recognise the importance of scientific quantities and understand how they are determined.</p> <p>WS 4.3 Use SI units (eg kg, g, mg; km, m, mm; kJ, J) and IUPAC chemical nomenclature unless inappropriate.</p> <p>WS 4.4 Use prefixes and powers of ten for orders of magnitude (eg tera, giga, mega, kilo, centi, milli, micro and nano).</p> <p>WS 4.5 Interconvert units.</p> <p>WS 4.6 Use an appropriate number of significant figures in calculation.</p> <p>MS 1a Recognise and use expressions in decimal form.</p> <p>MS 1b Recognise and use expressions in standard form.</p> <p>MS 1c Use ratios, fractions and percentages.</p>	<ul style="list-style-type: none"><li>• Understand the use of the multipliers in equations in normal script before a formula and in subscript within a formula.</li><li>• Calculate the percentage by mass in a compound given the relative formula mass and the relative atomic masses.</li><li>• Explain any observed changes in mass in non-enclosed systems during a chemical reaction given the balanced symbol equation for the reaction and explain these changes in terms of the particle model.</li><li>• Represent the distribution of results and make estimations of uncertainty</li><li>• Use the range of a set of measurements about the mean as a measure of uncertainty.</li><li>• Understand that the measurement of amounts in moles can apply to atoms, molecules, ions, electrons, formulae and equations, for example that in one mole of carbon (C) the number of atoms is the same as the number of molecules in one mole of carbon dioxide (CO<sub>2</sub>).</li><li>• Use the relative formula mass of a substance to calculate the number of moles in a given mass of that substance and vice versa.</li><li>• Calculate the masses of substances shown in a balanced symbol equation</li><li>• Calculate the masses of reactants and products from the balanced symbol equation and the mass of a given reactant or product.</li><li>• Balance an equation given the masses of reactants and products.</li><li>• Students should be able to change the subject of a mathematical equation.</li><li>• Explain the effect of a limiting quantity of a reactant on the amount of products it is possible to obtain in terms of amounts in moles or masses in grams.</li></ul>	<p>KS3 Y8 Atoms, elements &amp; compounds</p> <p>Y8 Chemical reactions</p> <p>4.1 Atomic structure &amp; The periodic Table</p>	<p>GCSE: Y10 – 4.4.1 Reactivity of metals - extraction of metals</p> <p>Y10 - 4.4.2 Reactions of acids</p> <p>A-level: Y12 3.1.2 Amount of substance</p>
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